20 Acids, bases, and pH

Specification reference	Checklist questions	
5.1.3 a i	Can you describe a Brønsted–Lowry acid as a species that donates a proton and a Brønsted–Lowry base as a species that accepts a proton?	☐
5.1.3 a ii	Can you use the term *conjugate acid–base pairs*?	☐
5.1.3 a iii	Can you describe monobasic, dibasic, and tribasic acids?	☐
5.1.3 b ii	Can you describe the role of H^+ in the reactions of acids with metals and bases (including carbonates, metal oxides and alkalis), using ionic equations?	☐
5.1.3 c i	Can you describe the acid dissociation constant, K_a, for the extent of acid dissociation?	☐
5.1.3 c ii	Can you describe the relationship between K_a and pK_a?	☐
5.1.3 d	Can you use the expression for pH as: $pH = -\log[H^+]$ $[H^+] = 10^{-ph}$	☐
5.1.3 e	Can you use the expression for the ionic product of water, K_w?	☐
5.1.3 f i	Can you calculate pH for strong monobasic acids?	☐
5.1.3 f ii	Can you calculate pH for strong bases using K_w?	☐
5.1.3 g	Can you calculate pH, K_a or related quantities, for a weak monobasic acid using approximations	☐

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Specification reference	Checklist questions	
5.1.3 h	Can you describe the limitations of using approximations to K_a related calculations for 'stronger' weak acids?	☐